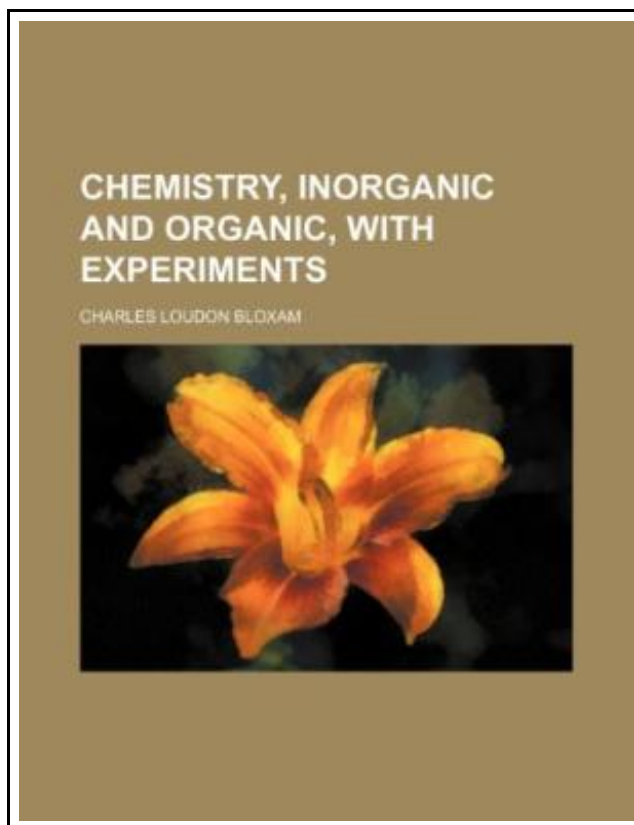


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1895 Excerpt: . easily imparts oxygen to other substances; sulphur, mixed with six times its weight of PbO., may be ignited by friction; hence this oxide is a common ingredient in lucifer-match compositions. Its oxidising property is frequently turned to account in the laboratory; for example, in absorbing sulphur dioxide from gaseous mixtures by converting it into sulphate of lead;  $PbO + SO_2 = PbSO_4$ . Dioxide of lead is not dissolved by dilute acids, and has no basic properties, although certain salts of the type  $PbX_4$  are known; it is even sometimes called plumbic anhydride, for it acts upon potassium hydroxide, yielding potassium plumbate ( $K_2PbO_4$ ), which has been crystallised from an alkaline solution, but is decomposed by pure water. Lead dioxide evolves  $O_2$  from  $HCl$  when heated, and gives, at first, a brown solution (containing  $PbCl_4$ ) which yields a brown precipitate with ammonia, but if the solution be boiled till all the  $Cl$  is expelled, it becomes colourless  $PbCl_2$ , and gives a white precipitate with ammonia.  $PbO$  is converted into  $PbO_2$  by ozone and by hydrogen peroxide. Lead hydroxide,  $Pb(OH)_2$ , has not been obtained, but  $Pb(OH)_2 \cdot PbO$  is formed as a white precipitate when air and water, free from  $CO_2$ , attack lead. The same substance is precipitated by alkalies from solutions of lead salts. It becomes  $PbO$  when heated to  $145^\circ C$ . The compound  $Pb(OH)_2 \cdot 2PbO$  crystallises in octahedra from a solution of basic lead acetate mixed with ammonia. Lead nitrate,  $Pb(NO_3)_2$ , crystallises in white octahedra from a solution...

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